



## AN ELECTROCHEMICAL CHARACTERISTICS OF PROMETHAZIE HCL USING IONSELECTIVE ELECTRODES

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### ARTICLE INFO

### ABSTRACT

#### Key Words

Promethazine HCl., sensor and ionophore.



The highly selective and sensitive PVC based membrane sensor was fabricated by using 1,3,5-tris [(2,3-dihydroxybenzylamino) amino methyl]cyclohexane (L) as a neutral membrane carrier for the potentiometric determination of Promethazine HCl. The membrane with the composition of 33% PVC, 64% Plasticizer, 3% PM-PMD(Ionophore) and 1% NaTPB was found to be best in terms of response characteristics of sensor assembly. The proposed membrane sensor has very low detection limit of  $2.0 \times 10^{-8}$  M, within the concentration range of  $5.0 \times 10^{-8}$  –  $1.0 \times 10^{-2}$  M, and has fast response time of about 10s. The proposed sensor was used for the selective determination of Promethazine HCl. in different synthetic as well as real sample.

## INTRODUCTION

Antihistamine (Histamine antagonist) is a pharmaceutical drug that inhibits action of histamine by blocking it from attaching to histamine receptors. There are several types of antihistamine drugs.  $H_1$  antihistamines are used to treat symptoms of allergy, such as runny nose and watery eyes,  $H_2$  antagonists (cimetidine), which are widely used for the treatment of acid reflux and stomach ulcers, because they decrease gastric acid production [1]. The  $H_3$  and  $H_4$  do not yet have a defined clinical use, although a number of drugs are currently in human trials [2].

The increasing use of ion sensors in the fields of environmental, agricultural and medicinal analysis is stimulating analytical chemists to develop new sensors for the fast, accurate, reproducible and selective determination of various species. In the past few decades, considerable efforts have led to the development of selective sensors for various medicinal compounds [3]. Histamine is derived from the decarboxylation histidine (amino acid). The decarboxylation is catalyzed by an enzyme histaminase, which is also involved in the metabolism of the bioactive amines. They cause the tissues in our nose to swell, our nose and eyes to run and our eyes, nose and sometime mouth to itch. Sometime they also cause itchy rash on our skin, called hives .

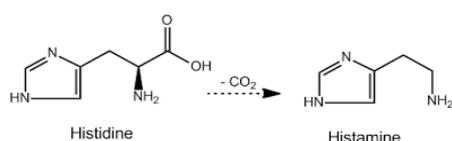
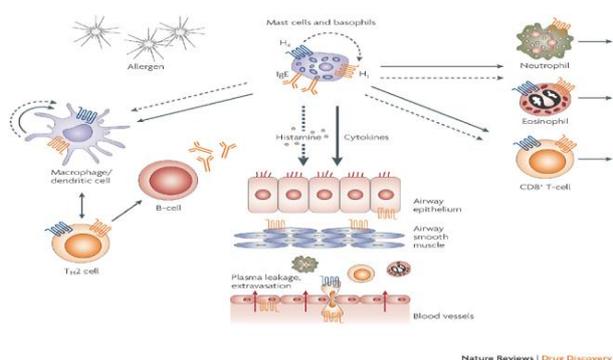


Fig:1 Decarboxylation of histidine to histamine

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**Fig: 2** Mechanism of action of antihistaminic drugs



## ION SELECTIVE ELECTRODES:

An ion selective electrode (ISE) measure the activity of an ion in a solution by measuring the electric potential formed across a membrane when the electrode is submerged in the solution. In order to measure the electrode potential developed at the ion-selective membrane the ISE/pH electrode must be immersed in the test solution together with a separate reference system and the two must be connected via a millivolt measuring system. At equilibrium, the electrons added or removed from the solution by the ISE membrane (depending on whether it is cation or anion sensitive) are balanced by an equal and opposite charge at the

reference interface. This causes a positive or negative deviation from the original stable reference voltage, which is registered on the external measuring system [4]. The relationship between the ionic concentration (activity) and the electrode potential is given by the Nernst equation:

$$E = E^0 + (2.303RT/nF) \times \text{Log}(A)$$

**Fig: 3** Ion Selective electrodes in pharmaceutical analysis

## EXPERIMENTAL:

(Determination of Promethazine HCl)

**Reagents and Equipments :** All the reagents of analytical grade and were used as received. High molecular weight polyvinyl chloride (PVC), Ammonium phosphomolybdate, Promethazine HCl, DBP, DBBP, OA, CN, THF, Ion analyzer, pH meter and Saturated calomel electrodes.

## Preparation of ion pair- compound

### Steps: 1 Ion-Pair Formation

Ion-pair compound of Promethazine HCl 20ml was mixed with 20ml of Amm. Phosphomolidbdate (at room for temp. 1hr) Precipitate of (PM-PMD) was obtained after one hour, and then precipitates was filtered off wash with water and dried.

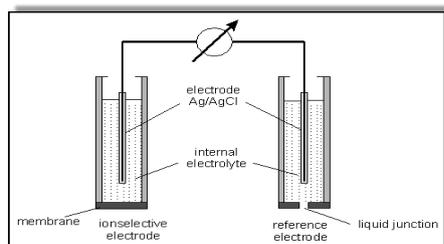
### Step: 2 Promethazine Selective PVC Membrane Electrode (Ionophore)

To prepare membrane electrode the membrane components i.e PM-PMD: Plasticizers: PVC in the ratio of (3: 64: 33) was (dissolved in THF) Resulting solution transferred into glass dish of 2cm diameter. The solvent was allowed to evaporate until conc. Mixture was obtained. A Pyrex tube dipped into conc. Mixture for 10sec. A transparent membrane about 0.3mm thickness was formed. Glass then pulled out and kept at room temperature for 5 hr. Tube was filled

with an internal solution of 0.001M Promethazine HCl solution. The electrode was conditioned for 24h by soaking in a  $1.0 \times 10^{-2}$ M of Promethazine solution

**Step: 3** Cell assembly for potential measurements:

Ag/AgCl, 0.1M |  
Internal reference solution 0.001M | test solution | 1M KCl, (KCl) Ag/AgCl



**Fig.4** Cell Assembly

## RESULT AND DISCUSSION:

The membrane electrode (no. 1) based on DBP as plasticizer has a detection limit of  $1.0 \times 10^{-6}$  M in a linear working concentration range of  $1.0 \times 10^{-6}$  M –  $1.0 \times 10^{-1}$  with slope of  $50.5 \pm 0.3$  (mV/dec. of activity). It was observed that 62 – 65% of the plasticizer as membrane components gives the best possible response. It was observed that the ionophore more than 3% (w/w) as membrane component does not improved the detection limit and linear concentration range. pH effect: It was observed that the potential of electrode assembly remains almost same in a pH range of 2.5 to 6.0.

**CONCLUSION:** A promethazine phosphomolibdate (PM-PMD) ion-pair compound was used as electroactive material for construction of promethazine selective electrode. The electrode of the composition of PVC: PM-PMD: DBP of 33: 3: 64 (% w/w) has a detection limit of  $1.0 \times 10^{-6}$  M in a linear concentration range of  $1.0 \times 10^{-6}$  –  $1.0 \times 10^{-1}$  M with a slope of calibration curve of  $50.5 \pm 0.3$  (mV/decay of activity). The electrode can be used in a pH range of 2.5 – 6.0 for a period of 4

weeks and has fast response time of about 5 s. The selectivity coefficient calculated by MPM method indicates that the electrode can be allied for the determination of promethazine in presence of other interfering ions.

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